

Ionic bonding - occurs b/w m + nm (give/take  $e^-$ )  
 Covalent bonding - occurs b/w 2 + nm (share  $e^-$ )  
 Metallic bonding - occurs b/w 2 + metals ( ? )

## Ionic Bonding

Force of attraction b/w oppositely charged particles

↳ ions = atoms that have a charge (+/-)

$\#p^+ \neq \#e^-$

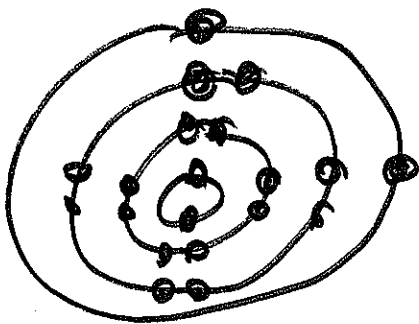
cation = +, metals (give  $e^-$ )

charge = + group #  $\text{Ca} \rightarrow \text{gives } e^- \rightarrow \text{Ca}^{+2}$

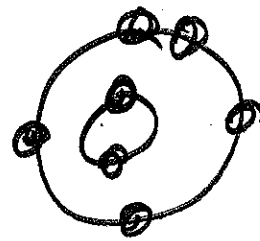
anion = -, nonmetals (take  $e^-$ )

charge = group # - 8

## Give or Take e<sup>-</sup>



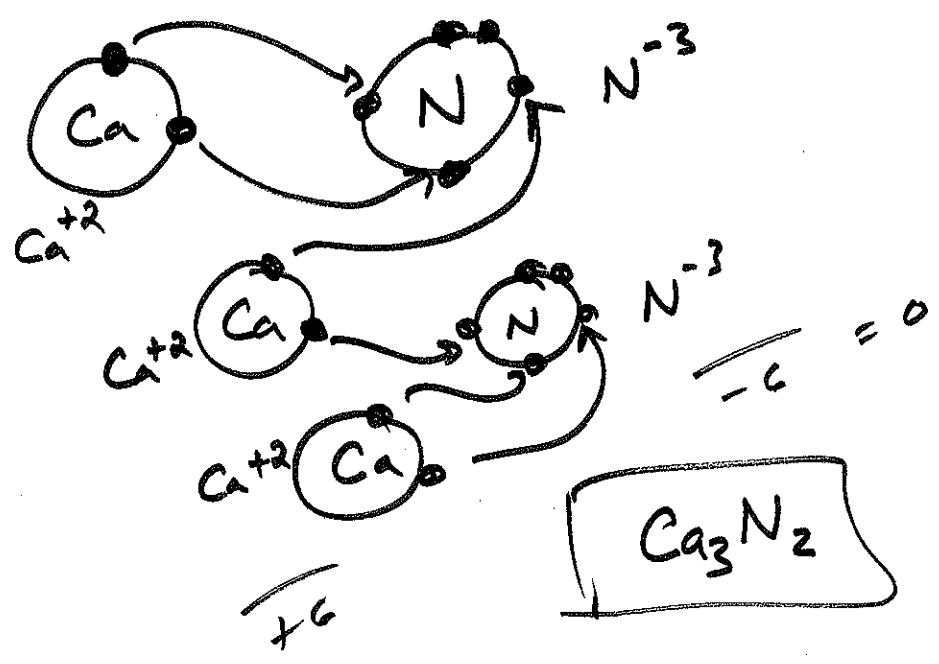
give 2,  $\text{Ca}^{+2}$



take 3,  $\text{N}^{-3}$

# Drawing Compounds WS (only valence shell)

Calcium and nitrogen



## Writing Ionic Formulas WS

